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# Field Testing a Pilot-Scale System for Alkaline Dehydration of Source-Separated Human Urine: A Case Study in Finland

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Alkaline dehydration can treat human urine to produce a dry and nutrient-rich fertilizer. To evaluate the technology at pilot-scale, we built a prototype with capacity to treat 30 L urine d<sup>-1</sup> and field tested it for the first time at a military training camp in Finland. We operated the system for 3 months and monitored the recovery of nutrients, end-product composition, physicochemical properties and energy consumption. Results revealed that the system received less urine than anticipated, but achieved high dehydration rates (30–40 L d<sup>-1</sup>m<sup>-2</sup>), recovered 30 ± 6% N, and yielded end-products with 1.4% N, 0.9% P, and 8.3% K. However, we demonstrated that the system had potential to recover nearly 70% N and produce fertilizers containing 13.2% N, 2.3% P, and 6.0% K, if it was operated at the design capacity. The energy demand for dehydrating urine was not optimized, but we suggested several ways of reducing it. We also discussed concerns surrounding non-essential heavy metals, salts, and micropollutants, and how they can be overcome to safely recycle urine. Lastly, we pointed out the research gaps that need to be addressed before the technology can be implemented at larger scale.

**Keywords:** circular bioeconomy, decentralized systems, nutrient recycling, resource recovery, sanitation, source-separation, urine drying, wastewater treatment

## INTRODUCTION

In the field of wastewater management, the idea of separately collecting and treating different fractions of wastewater at the point where it is produced is rather innovative (Larsen et al., 2013). A source-separating sanitation system treats various sub-streams of domestic wastewater according to their composition (Friedler et al., 2013). Such an approach allows resources present in wastewater to be recovered for reuse or recycling (e.g., plant-essential nutrients, water and heat). It can also prevent the dispersion of potentially harmful substances in the environment by removing them as close as possible to the source (e.g., micropollutants).

Although the idea of source separation has been around for a while, it has struggled to move from concept to practice (McConville et al., 2017). One of the hurdles to its implementation lies in developing technologies to recycle human urine, wastewater's most nutrient-rich fraction (Winker et al., 2009). Several research groups across the world have been working on this niche topic

and have investigated different options for on-site urine treatment, many of which are reviewed in Harder et al. (2019). At the Swedish University of Agricultural Sciences we have opted for an alternative approach, one where fresh urine is dehydrated to concentrate nutrients within an alkaline medium, resulting in the production of a dry fertilizer. This approach, which we call alkaline urine dehydration, has been studied extensively in the laboratory (Dutta and Vinneras, 2016; Senecal and Vinneras, 2017; Senecal et al., 2018; Simha et al., 2018, 2020a).

In early 2019, we built a pilot-scale dry sanitation system with capacity to treat 30 L urine  $\text{d}^{-1}$ . The system was designed to add fresh human urine continuously to a mixture of wood ash and calcium hydroxide (to prevent enzymatic ureolysis) and to concentrate the urine by hot air convective dehydration. It was operated and monitored for a period of 3 months, the results of which are summarized in this paper. To assess whether alkaline dehydration is a suitable on-site urine treatment technology, we focused on the following aspects – (i) the recovery of nutrients, particularly nitrogen; (ii) the elemental composition of the end-product; (iii) the concentrations of micropollutants and heavy metals; and, (iv) the energy demand for dehydrating urine. We also highlighted aspects that need further research, before the technology can be practically implemented.

## Piloting Context

The pilot-scale study was carried out at a Finnish Army training site (Camp Mauri) belonging to the Pori Brigade, which is garrisoned in Säkylä, Finland. The brigade trains more than 4000 conscripts every year and supports voluntary training exercises and courses for the Finnish Defence Forces' reserves. A lot of the drills take place in forest terrain and specifically designed sites at the garrison area. Camp Mauri represents the latter and is used by people undergoing training and thus, it needs to provide adequate sanitation facilities. The camps normally have permanent toilets but may temporarily rent field toilets such as those used at events or use disposable box toilets if they are located in areas that are hard to access. However, due to the 2016 ban in Finland on the landfilling of biodegradable waste, sewage from the camps needs to be collected and transported to a nearby wastewater treatment plant (Ministry of the Environment, 2016). One way to reduce the need for sewage transportation from off-the-grid sites is to concentrate the urine at the point of collection (Simha et al., 2020b). The present study explored this possibility by implementing urine source-separation and on-site treatment with alkaline dehydration.

## MATERIALS AND METHODS

### Alkaline Media

For preparing the dehydration media, we used a mixture of wood ash and calcium hydroxide. The ash was collected from a cattle farm burning wood chips for bioenergy production (Lövsta Kött AB, Sweden) and pre-treated in a furnace (LH30/12, Nabertherm GmbH, Germany) for 6 h at 550°C, as described in Simha et al. (2018). Calcium hydroxide ( $\text{Ca}(\text{OH})_2$ ) was supplied

by Nordkalk Corporation, Sweden and used without any pre-treatment. The media contained 18.75%  $\text{Ca}(\text{OH})_2$  and 81.25% pre-treated wood ash by weight.

### Pilot Design and Setup

An on-site urine collection and alkaline dehydration system was developed and installed at a training camp of the Pori Brigade, an Army Unit of the Finnish Defence Forces located in Säkylä, Finland. The pilot-scale setup (Figure 1) comprised of a modified portable dry sanitation system (Biomaja Nordic Oy, Finland), an alkaline urine dehydrator and an industrial dehumidifier (DSC-70ES, Woods AB, Sweden).

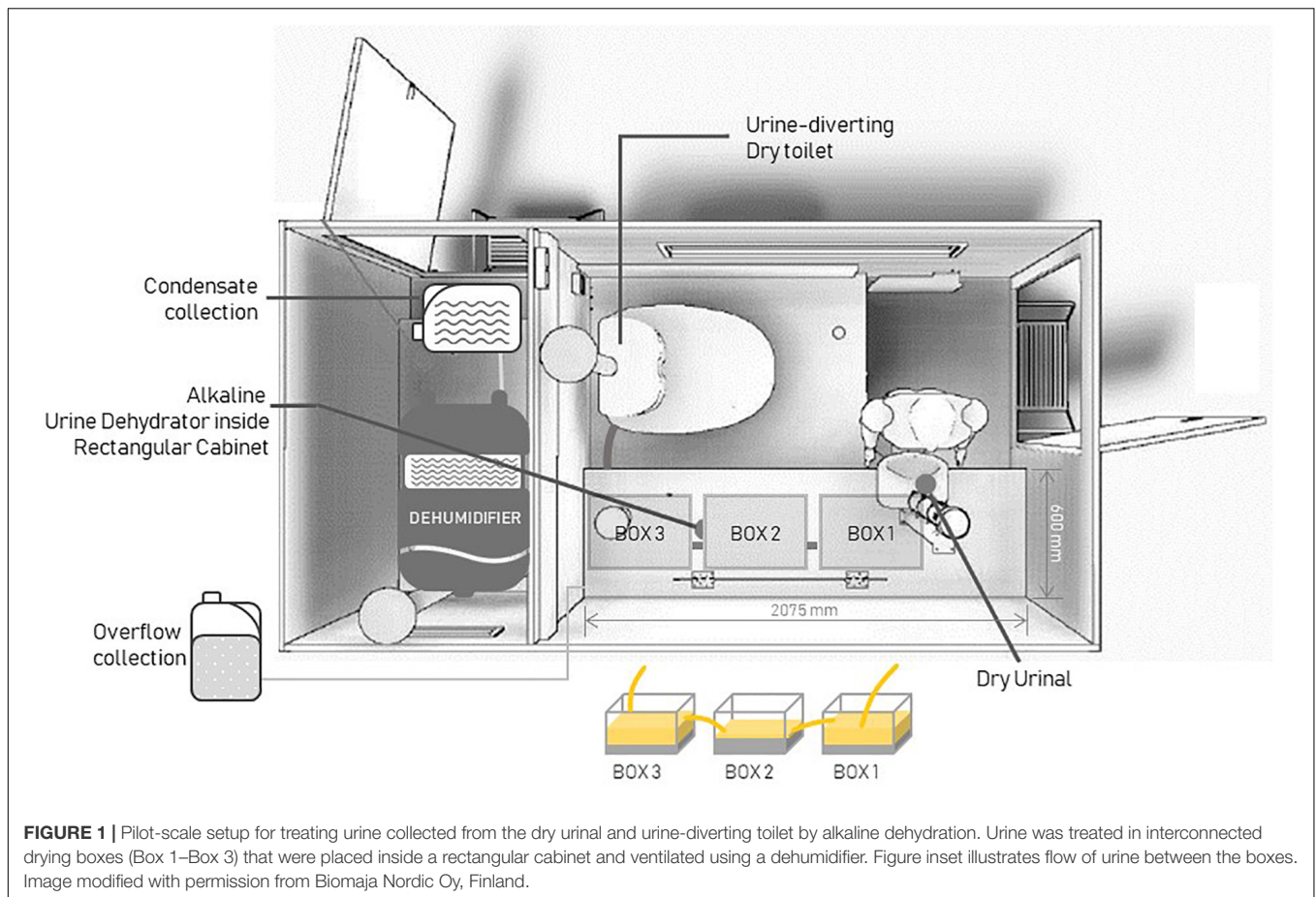
The dry sanitation system comprised of two rooms. The larger room (2075 × 1600 × 2200 mm) had a dry urinal and a dry toilet (Biolan Oy, Finland) that was retrofitted using a Privy 501 kit (Separett AB, Sweden) to source-separate urine. The dehumidifier was placed in the smaller room (1000 × 1600 × 2200 mm) and provided airflow (450  $\text{m}^3 \text{h}^{-1}$  according to the manufacturer) throughout the system. The power rating of the dehumidifier was 630 W  $\text{day}^{-1}$  when operating at 20°C and 70% relative humidity.

A rectangular cabinet (2075 × 600 × 550 mm) located below the urinal that usually contains a urine collection tank was repurposed as an on-site urine dehydrator. In this cabinet, three polypropylene boxes (580 × 450 × 300 mm; Klämtare, IKEA of Sweden) were placed in a series configuration and interconnected with short (<200 mm) polypropylene pipes ( $\text{Ø}$  50 mm). At the start of the treatment, all the boxes contained the drying media, a mixture of wood ash and  $\text{Ca}(\text{OH})_2$ . The boxes were covered with acrylic reinforced lids and each lid housed two heater fans (40–137, Biltema, Sweden) that provided air flow directly onto the surface of the drying media. The fans were fitted with a ceramic heating element and operated on two different settings that had a power rating of either 700 W or 1200 W.

A short pipe (<100 mm) connected Box 1 to the urinal while Box 3 was connected to the diverting toilet with a flexible hose (<500 mm). Urine entered Box 2 only when either or both the other boxes overflowed. It was important that the drying boxes were placed close to the point of urine collection since biofilms usually present in the piping network of larger collection systems can degrade urea into ammonia before the urine is alkalinized to inhibit ureolysis. A 25 L jerry can containing 1 kg wood ash was placed outside the system and connected to the Box 3 to collect urine in case of a system overflow. For airflow, holes ( $\text{Ø}$  80 mm) were drilled into the sides of the rectangular cabinet – the first served as the ambient air inlet to the dehydrator while the second interconnected the dry sanitation system's two rooms and thus, the dehydrator with the dehumidifier. The dehumidifier condensate, if any, was pumped into a 25 L jerry can.

### Pilot Operation

The pilot-scale system was designed to dry fresh urine collected from approximately 100 toilet users every day (or a maximum of 30 L urine  $\text{day}^{-1}$ ), for a period of 3 months between March and May 2019. The piloting was divided into three 1-month treatment periods. At the start of a treatment period, 6.5 kg wood ash and 1.5 kg  $\text{Ca}(\text{OH})_2$  were added to each of the dehydrator's three



boxes, thoroughly mixed, and wetted with approximately 20 kg water box<sup>-1</sup> (to prevent air suspension of the media).

The system was then switched on and allowed to dry out the water (around 1–2 days), after which the pilot operation was handed over to the Construction Establishment of Finnish Defence Administration. Over the month, the system was used by military personnel training at the camp site. At the end of the treatment, the system was switched off, the drying boxes were collected, emptied, and the end-product was sampled.

## Monitoring and Analysis

Samples were taken for the drying media, the end-products, fresh urine as well as the dehumidifier condensate. Fresh urine could only be collected once during the pilot since it was not certain when the military personnel were at the site and using the sanitation system. The composition of the urine is shown in **Table 1**. The dehumidifier condensate was sampled if it was available during the site visits.

For sampling the end-product after every treatment period, the drying boxes were detached from the dehydrator and their contents were thoroughly mixed. For each box, composite samples were collected by blending grab samples taken from at least five random locations in the box. At least three composite samples were taken per box for every treatment period.

For all the collected samples, we analyzed for Total Solids (TS), Volatile Solids (VS), pH<sub>1:5</sub> and EC<sub>1:5</sub> as described in Simha et al. (2020a). We also analyzed for the concentrations of different elements and micropollutants – in total, 19 elements, 3 anions, 126 pharmaceuticals and 8 hormones were analyzed. The methods used for the various analyses are described in the **Supplementary Information Tables S1, S2**.

For monitoring air temperature and relative humidity in the system, data loggers (set to 5 min intervals) with inbuilt sensors (Tinytag Ultra 2 TGU-4500, Intab, United Kingdom) were placed at seven different positions in the system – ambient, dehydrator air inlet, dehumidifier inlet and outlet, and inside the three drying boxes (**Supplementary Information, Supplementary Figure S1**). The electricity consumption of the heater fans and the dehumidifier was monitored using plug-in meters (EMT707CTL, Clas Ohlson, United Kingdom).

## Ethics Review

The ethics review of this study was carried out by the Ethics Committee of the Tampere Region, which handles the ethical reviews of proposed non-medical research. It operates under the guidelines issued by the Finnish National Board on Research Integrity (TENK). According to the review issued on 18 February 2019 (6/2019), the present study did not include ethical issues

**TABLE 1** | Measured composition of urine ( $n = 3$ ) collected at the study site.

Measurement	Unit	Average $\pm$ Std. Dev
Total N	[mg L <sup>-1</sup> ]	8600 $\pm$ 86
Total P	[mg L <sup>-1</sup> ]	637 $\pm$ 5
K	[mg L <sup>-1</sup> ]	1467 $\pm$ 47
Ca	[mg L <sup>-1</sup> ]	96 $\pm$ 0.5
Mg	[mg L <sup>-1</sup> ]	62 $\pm$ 0.8
Na	[mg L <sup>-1</sup> ]	2533 $\pm$ 47
As	[ $\mu$ g L <sup>-1</sup> ]	19 $\pm$ 0
B	[ $\mu$ g L <sup>-1</sup> ]	653 $\pm$ 5
Cd	[ $\mu$ g L <sup>-1</sup> ]	0.3 $\pm$ 0.1
Co	[ $\mu$ g L <sup>-1</sup> ]	0.5 $\pm$ 0.01
Cr	[ $\mu$ g L <sup>-1</sup> ]	8 $\pm$ 0.2
Cu	[ $\mu$ g L <sup>-1</sup> ]	25 $\pm$ 9
Fe	[ $\mu$ g L <sup>-1</sup> ]	87 $\pm$ 17
Mn	[ $\mu$ g L <sup>-1</sup> ]	21 $\pm$ 0.5
Mo	[ $\mu$ g L <sup>-1</sup> ]	45 $\pm$ 0.5
Ni	[ $\mu$ g L <sup>-1</sup> ]	2.8 $\pm$ 0.1
Pb	[ $\mu$ g L <sup>-1</sup> ]	0.6 $\pm$ 0.1
Se	[ $\mu$ g L <sup>-1</sup> ]	35.3 $\pm$ 1
Zn	[ $\mu$ g L <sup>-1</sup> ]	303 $\pm$ 5
Cl <sup>-</sup>	[mg L <sup>-1</sup> ]	3967 $\pm$ 125
SO <sub>4</sub> <sup>2-</sup>	[mg L <sup>-1</sup> ]	1200 $\pm$ 0
F <sup>-</sup>	[mg L <sup>-1</sup> ]	61 $\pm$ 7

that would prevent its implementation as planned in the research proposal.

## Calculations

Treating urine by alkaline dehydration reduced its mass, the extent of which was calculated using Eq. 1.

$$mass.cf_{WB} = \left( \frac{m_{media} + m_{urine}}{m_{end-product}} \right), \quad (1)$$

where  $mass.cf_{WB}$  is the mass concentration factor on wet basis and  $m_{media}$ ,  $m_{urine}$ , and  $m_{end-product}$  are the weight of the drying media at the start of the treatment, the total urine added during the treatment, and the end-product, respectively.

To estimate the total urine inflow to the system ( $m_{urine}$ ), we performed a mass balance for P since there is complete recovery of P from urine during alkaline dehydration (Simha et al., 2020a). We calculated urine's contribution of P ( $m_{P,urine}$ ) to the total P recovered in the end-product ( $m_{P,end-product}$ ) using Eq. 2, where  $m_{P,media}$  is the P content of the drying media at the start of the treatment.

$$m_{P,urine} = m_{P,end-product} - m_{P,media} \quad (2)$$

$$m_{urine} = \frac{m_{P,urine}}{C_{P,urine}} \quad (3)$$

The P content of the sampled urine ( $C_{P,urine}$ ) was measured as described in the earlier section. We also estimated the N content (Eq. 4) and P content (Eq. 5) of urine, as suggested by Jönsson and Vinnerås (2004). The data for total ( $m_{total\ protein}$ )

and vegetal protein ( $m_{vegetal\ protein}$ ) consumption of the system users was collected from the local canteen (See **Supplementary Information, Supplementary Table S3**) and the latest statistics for food supplied in Finland (FAO, 2017).

$$m_{N,urine+faeces} = 0.13 \times m_{total\ protein} \quad (4)$$

$$m_{P,urine+faeces} = 0.011 \times (m_{total\ protein} + m_{vegetal\ protein}) \quad (5)$$

Following Simha et al. (2020a), we used Eq. 6 to determine the recovery of N in the form of urea:

$$urea.rec_i = (urea.add_i + urea.rec_{i-1}) \times \left( 1 - \frac{t}{2 \times h_{1/2}} \right), \quad (6)$$

where,  $i$  is the time period,  $urea.add$  and  $urea.rec$  are the amounts of urea added and recovered, respectively. The half-life ( $h_{1/2}$ ) of urea at different temperatures was estimated as:  $1.352 \times 10^6 \times e^{-(0.1257T)}$ . We assumed that, of all the N excreted in urine, only urea-N (85% of the total N) can be recovered by alkaline dehydration, while the remaining is lost.

## Statistics

The data was tested for homogeneity of variances using the  $F$ -test and the Shapiro–Wilk test for normality. To identify significant differences between the drying boxes, either unpaired two-sample Student's  $t$ -test (for homoscedastic data) or Welch's  $t$ -test (for heteroscedastic data) was performed. The level of significance was set at  $p < 0.05$ . One-way analysis of variance was carried out to identify statistically significant differences in the urine treatment across the different treatment periods. When a significant difference was found, Tukey's Honest Significant Difference (HSD) test at 95% confidence level was done.

## RESULTS

### Mass Balance

The design capacity of the pilot was 900 L month<sup>-1</sup> but the volume of urine it received was considerably less. Based on the mass balance of P (Eq. 2, 3) and the analysis of the sampled urine ( $C_{P,urine} = 0.64\text{ g}_P\text{ L}^{-1}$ ), the incoming urine volume was estimated to be approximately 35, 62, and 90 L in March, April, and May, respectively (**Supplementary Information, Supplementary Figure S2**). However, if the composition of urine is estimated using the Jönsson and Vinnerås (2004) approach as 11.4 g<sub>N</sub> L<sup>-1</sup> and 1.12 g<sub>P</sub> L<sup>-1</sup>, the urine volume treated by the pilot would have been 20, 25, and 51 L in March, April, and May, respectively. The three boxes of the alkaline dehydrator also received different amounts of urine – no urine entered Box 2 throughout the pilot testing, while Box 1 received 46  $\pm$  22 L month<sup>-1</sup> and Box 3 received 16  $\pm$  7 L month<sup>-1</sup>.

The end-product from Box 1 weighed 13.3  $\pm$  3.7 kg, while that from Box 3 weighed 10.5  $\pm$  2 kg. The mass concentration factor revealed that the end-products weighed 3–5 times less than the added urine and drying substrate. The dehumidifier condensate was available only at the end of March and May, where the amounts collected were 2.4 and 22.2 kg, respectively.

## Nitrogen Recovery

Based on the estimated urine volumes, the mass balance revealed that the recovery of N in the end-product was 29, 31 and 38% for March, April and May (Supplementary Information, Supplementary Figure S2). The recovery was lower if the N content of urine was estimated using the Jönsson and Vinnerås (2004) method. The expected recovery of N was 66%, while the expected urea-N recovery at this temperature is 78% (Eq. 6). This was calculated using the half-life of urea assuming a constant drying temperature of 60°C.

The two drying boxes received different amounts of urine and recovered different amounts of N ( $34 \pm 5\%$  in Box 1 and  $27 \pm 4\%$  in Box 3), but the differences were not statistically significant ( $p > 0.05$ ). For Box 1, the recovery of N was significantly different over the three treatment periods, yet this was not the case for Box 3.

The dehumidifier condensate also captured N – the concentration of  $\text{NH}_4\text{-N}$  in the condensate was  $194 \pm 13 \text{ mg}_\text{N} \text{ L}^{-1}$  and the total-N concentration was  $511 \pm 132 \text{ mg}_\text{N} \text{ L}^{-1}$ . However, in the calculations for overall N recovery of the pilot system, the N captured in the condensate was disregarded.

## End-Product Composition

The N-P-K composition (as% of TS) of the end-products collected in March, April and May was 0.5–0.7–6.1, 0.8–0.8–6.8, and 1.4–0.8–7.4, respectively (Table 2). The maximum N content was 2.1%, for the end-product collected from Box 1 at the end of May. On average, the end-products contained 17% Ca, 2.8% Mg, 1% Mn and 0.4% S. The average  $\text{Na}^+$  and  $\text{Cl}^-$  concentrations were 0.9 and 1.3%, respectively. Heavy metals Cu and Zn were also detected –  $1946 \pm 320 \text{ mg kg}_\text{p}^{-1}$  and  $3844 \pm 523 \text{ mg kg}_\text{p}^{-1}$ ,

**TABLE 2 |** The mass, total solids (TS), total solids adjusted for loss of urea ( $\text{TS}_{\text{adj}}$ ) and elemental composition of the end-products collected in the three treatment periods.

Property	March	April	May
Mass [kg]	18.9	22.8	29.7
TS [kg]	18	19.2	20
$\text{TS}_{\text{adj}}$ [kg]	18.2	19.5	20.5
<b>Elemental composition [as% of <math>\text{TS}_{\text{adj}}</math>]</b>			
N	0.5 <sup>C</sup>	0.8 <sup>B</sup>	1.4 <sup>A</sup>
P	0.7 <sup>B</sup>	0.8 <sup>AB</sup>	0.8 <sup>A</sup>
K	6.1 <sup>A</sup>	6.8 <sup>A</sup>	7.4 <sup>A</sup>
Ca	19.2 <sup>A</sup>	17.3 <sup>B</sup>	15.0 <sup>C</sup>
Na	0.6 <sup>C</sup>	1.2 <sup>A</sup>	0.9 <sup>B</sup>
Mn	0.9 <sup>B</sup>	0.8 <sup>C</sup>	1.0 <sup>A</sup>
S	0.3 <sup>C</sup>	0.5 <sup>A</sup>	0.4 <sup>B</sup>
Mg	2.9 <sup>A</sup>	2.6 <sup>B</sup>	3.0 <sup>A</sup>
Cu	0.009 <sup>B</sup>	0.009 <sup>B</sup>	0.011 <sup>A</sup>
Zn	0.019 <sup>B</sup>	0.019 <sup>B</sup>	0.022 <sup>A</sup>
Cd	0.0004 <sup>A</sup>	0.0004 <sup>A</sup>	0.0004 <sup>A</sup>
$\text{Cl}^-$	0.8 <sup>C</sup>	1.3 <sup>B</sup>	1.8 <sup>A</sup>

Values marked with the same letters show there is no significant difference ( $p < 0.05$ ).

respectively. The average Cd content of the end-product was  $529 \pm 26 \text{ mg kg}_\text{p}^{-1}$ .

## Micropollutants

The concentrations of pharmaceuticals and hormones detected in fresh urine and the end-product collected from the second treatment period (April) is presented in Table 3. In the untreated urine, only 8 out of the 126 monitored pharmaceuticals and 4 out of the 8 hormones were detected. This included the following therapeutic groups: non-steroidal anti-inflammatory drugs (ibuprofen, naproxen),  $\beta$ -blockers (bisoprolol, metoprolol), analgesics (acetaminophen), CNS stimulants (caffeine), glucocorticosteroids (hydrocortisone), a decongestant (xylometazoline) and two antibiotics (tetracycline, doxycycline). After alkaline dehydration treatment, none of the hormones were detected in the solid end-product. However, the pharmaceuticals bisoprolol, metoprolol, ibuprofen, caffeine, xylometazoline and naproxen were detected. The pharmaceuticals partitioned primarily to the solid phase, as only caffeine ( $24 \mu\text{g L}^{-1}$ ) was detected in the dehumidifier condensate.

## Physicochemical Properties

The drying substrate had an average  $\text{pH}_{1:5}$  of  $13.3 \pm 0.02$  and  $\text{EC}_{1:5}$  of  $49 \pm 0.7 \text{ mS cm}^{-1}$  at the start of the treatment. All through the pilot, the end-products had  $\text{pH}_{1:5} > 10$ , which is a pre-requisite to prevent enzymatic ureolysis (Geinzer, 2017). Box 3 had significantly higher alkalinity ( $p < 0.001$ ) as well as electrical conductivity ( $p < 0.05$ ) than Box 1 for all the treatment periods (Table 4). The dehumidifier condensate had a  $\text{pH}_{1:5}$  of  $9.0 \pm 0.1$  and  $\text{EC}_{1:5}$  of  $801 \pm 5 \mu\text{S cm}^{-1}$ . The average dry matter content of the end-products collected from Box 1 and Box 3 was  $80 \pm 16\%$  and  $88 \pm 13\%$ , respectively, while that of the condensate water was 99.9%. The average air temperature differed significantly between the two boxes ( $p < 0.001$ ) but also between the different treatment periods ( $p < 0.05$ ). For long periods, there was large temperature overshoot above the desired average drying temperature of 60°C, reaching temperatures up to 90°C.

## Energy Demand

The pilot's total electric energy consumption was 1913, 2315, and 1185 kWh during March, April, and May, respectively. The dehumidifier consumed a considerable amount of energy ( $521 \pm 220 \text{ kWh month}^{-1}$ ), but we disregarded this when relating the total energy demand with the amount of urine treated by the pilot. The average electric energy demand for alkaline dehydration of urine was  $24.5 \pm 13 \text{ kWh kg}^{-1}$  urine. By using a conversion factor of 2.1 [for average European electricity mix; EC (2018)], the primary energy demand for the three treatment periods amounted to 76, 57, and 21 MJ  $\text{kg}^{-1}$  urine.

## DISCUSSION

### Nutrients

The idea with most technologies treating source-separated urine including alkaline dehydration is to recover resources, primarily

**TABLE 3** | Concentrations (mean  $\pm$  standard deviation) of pharmaceuticals and hormone compounds detected in urine ( $n = 3$ ) and in the end-product collected at the end of April ( $n = 2$ ) on dry matter (TS) basis.

Compound	CAS Number	Urine		End-product		
		Measured	LoD	Measured	Simulated	LoD
		$[\mu\text{g L}^{-1}]$	$[\mu\text{g L}^{-1}]$	$\text{mg kg}_{\text{TS}}^{-1}$	$\text{mg kg}_{\text{TS}}^{-1}$	$\text{mg kg}_{\text{TS}}^{-1}$
<b>Pharmaceuticals</b>						
Bisoprolol	66722–44–9	ND	1.00	0.010	–	0.005
Hydrocortisone	50–23–7	27.3 $\pm$ 0.9	1.00	<LoD	0.09	0.005
Ibuprofen	15687–27–1	118 $\pm$ 33	5.00	18.5 $\pm$ 5	0.38	0.01
Caffeine	58–08–2	3067 $\pm$ 47	1.00	3.1 $\pm$ 3.4	9.84	0.001
Xylometazoline	526–36–3	6.4 $\pm$ 0.3	0.10	0.9	0.02	NA
Naproxen	22204–53–1	7.4 $\pm$ 2.0	1.00	0.45 $\pm$ 0.36	0.02	0.2
Paracetamol	103–90–2	121 $\pm$ 20	2.00	<LoD	0.39	0.005
Tetracycline	60–54–8	12.3 $\pm$ 1.3	1.00	0.02 $\pm$ 0.02	0.04	0.005
Metoprolol	37350–58–6	ND	0.10	0.1 $\pm$ 0.13	–	0.005
Doxycycline	564–25–0	1.2 $\pm$ 0.1	1.00	<LoD	0.004	<0.005
<b>Hormones</b>						
Estradiol	50–28–2	1.3 $\pm$ 0.5	0.50	<LoD	0.004	0.05
Estriol	50–27–1	2.2 $\pm$ 1.0	0.50	<LoD	0.007	0.25
Estrogen	53–16–7	1.3 $\pm$ 0.3	0.50	<LoD	0.004	0.05
Testosterone	58–22–0	3.7 $\pm$ 2.6	0.10	<LoD	0.012	0.005

NA, Not Available. For the end-product, both the concentrations measured by sampling and those simulated by assuming that no degradation occurred during the treatment is presented along with the limit of detection (LoD).

**TABLE 4** | The physicochemical properties of the end-products collected in the two boxes at the end of the three treatment periods.

Box #	pH [-]	EC [ $\text{mS cm}^{-1}$ ]	TS [%]	VS [%]	T [ $^{\circ}\text{C}$ ]
<b>Box 1</b>					
March	13.1 $\pm$ 0.0 <sup>A,***</sup>	29.7 $\pm$ 1.3 <sup>A,*</sup>	96.7 $\pm$ 0.8 <sup>A,*</sup>	9.3 $\pm$ 0.6 <sup>B</sup>	40 $\pm$ 27 <sup>C,***</sup>
April	11.8 $\pm$ 0.2 <sup>C,***</sup>	22.8 $\pm$ 1.4 <sup>C,*</sup>	75.6 $\pm$ 1.7 <sup>B,***</sup>	10.7 $\pm$ 1.5 <sup>AB,*</sup>	66 $\pm$ 15 <sup>A,***</sup>
May	12.5 $\pm$ 0.0 <sup>B,***</sup>	26.1 $\pm$ 1.3 <sup>B,**</sup>	63.2 $\pm$ 2.8 <sup>C,***</sup>	12.1 $\pm$ 1.1 <sup>A,*</sup>	56 $\pm$ 15 <sup>B,***</sup>
<b>Box 3</b>					
March	13.4 $\pm$ 0.1 <sup>A</sup>	33.6 $\pm$ 2.7 <sup>A</sup>	94.1 $\pm$ 1.9 <sup>B</sup>	7.3 $\pm$ 2.1 <sup>B</sup>	36 $\pm$ 20 <sup>C</sup>
April	13.3 $\pm$ 0.1 <sup>B</sup>	27.3 $\pm$ 3.7 <sup>B</sup>	96.8 $\pm$ 0.3 <sup>A</sup>	8.7 $\pm$ 0.4 <sup>B</sup>	64 $\pm$ 14 <sup>A</sup>
May	13.1 $\pm$ 0.1 <sup>B</sup>	29.9 $\pm$ 1.2 <sup>AB</sup>	72.4 $\pm$ 0.9 <sup>C</sup>	10.7 $\pm$ 0.4 <sup>A</sup>	54 $\pm$ 15 <sup>B</sup>

Values are presented as mean  $\pm$  standard deviation ( $n = 6$ ). For each box, different letters indicate significant differences ( $p < 0.05$ ) row-wise across the treatment periods. For the same treatment period, values that differ significantly between the two boxes are indicated as: \* $p < 0.05$ , \*\* $p < 0.01$ , and \*\*\* $p < 0.001$ .

macronutrients. Our past studies (Senecal and Vinneras, 2017; Simha et al., 2018, 2020a) in the laboratory demonstrated the possibility of recovering almost all nutrients from urine, including N (>90%). This, however, was not the case in our pilot-scale setup where the recovery of N was only  $30 \pm 6\%$ , while we expected to recover 66% of the N at the design average treatment temperature of  $60^{\circ}\text{C}$ .

According to the thermodynamics of water evaporation, a large amount of heat (enthalpy of vaporization) is required for phase transition of liquid water to water vapor, approximately  $2358 \text{ kJ kg}^{-1}$  at  $60^{\circ}\text{C}$ . Water absorbs sensible heat from hot air to convert it into latent heat. This causes an evaporative cooling effect that reduces the air temperature. Our pilot system received <10 percent of the urine ( $<3 \text{ L d}^{-1}$ ) than what was intended. As a result, the evaporative cooling effect reduced and the temperature in the drying boxes overshot, reaching up

to  $90^{\circ}\text{C}$ . High temperature and high pH (>12) would have promoted chemical urea hydrolysis (Blakeley et al., 1982), which can explain the high N losses of the pilot system.

If the pilot system operated at the design capacity, both the N recovery as well as end-product composition would be very different. We estimate that drying 900 kg of urine with 24 kg of drying media (19.5 kg wood ash mixed with 4.5 kg  $\text{Ca}(\text{OH})_2$ ) will result in approximately 54 kg of end-product (Figure 2). The N-P-K value (as% of TS) of the final product should be 9.9–1.5–6.0, and it should also contain 0.8% S, 10.2% Ca, 1.6% Mg, 4.6% Na and 6.9%  $\text{Cl}^-$ . The end-product composition is similar to blended fertilizers (11–4.6–17.6) used in Finland and Sweden for vegetables, fruits and berries (Granngård, 2019; Yara, 2020). Blended mineral fertilizers can contain micronutrients such as Mg, S, B, Cu, Fe, Mn, Mo, and Zn, all of which are also present in urine and the end-product. Preliminary crop trials applying dried

	IN		+	=	OUT		+	Condensate ↓ & Losses ↑	
	MEDIA				URINE	END PRODUCT			
	g	% of TS			g	g			% of TS
Mass	24000		900000		54353				
TS	24000				51527				
Urea	0		14302		11128				
TS <sub>(urea adj)</sub>	24000				62654				
N	0	0	7740		5118.8	9.9			
P	198	0.8	573		771.2	1.5			
K	1773	7.4	1320		3093	6.0			
Ca	5160	21.5	86		5246	10.2			
Na	82	0.3	2280		2362	4.6			
Mn	274	1.1	0.02		274	0.5			
Cu	3	0.01	0.02		3	0.01			
Zn	6	0.03	0.3		6	0.01			
S	61	0.3	360		421	0.8			
Mg	817	3.4	0.9		818	1.6			
Cd	0.16	0.001	0.3		0.44	0.001			
Cl <sup>-</sup>	5	0.02	3570		3575	6.9			

**FIGURE 2 |** Simulated mass and nutrient balance (in g) for alkaline dehydration of 900 kg urine month<sup>-1</sup> using the pilot-scale system. The end-product composition is shown as a% of TS that was adjusted for loss of urea. The measured composition of urine shown in **Table 1** was used for the simulation and end-product mass was calculated using a *mass.cf<sub>WB</sub>* of 17 from Simha et al. (2020a).

urine have been promising (Martin et al., 2017) and sociological studies have indicated positive consumer acceptance (Simha et al., 2017; Brun, 2018).

Using a dehumidifier for venting out the exhaust air produced condensate water containing nearly 0.5 g<sub>N</sub> L<sup>-1</sup>, because the N loss from the drying boxes was high. Even if N losses can be limited, we can expect the condensate water to contain some N. One way to manage this water would be to soil infiltrate it, especially since it does not contain P (not detected, detection limit 0.01 mg L<sup>-1</sup>). This discharge would be within the Finnish regulations concerning the environmental loading of treated wastewater, even in groundwater sensitive areas (Ministry of the Environment, 2011).

## Energy

The energy demand of the pilot (21–76 MJ kg<sup>-1</sup> urine) was far more than what we expected, as operating it at the design capacity would consume 10–12 MJ kg<sup>-1</sup> urine; the calculation assumes that 30 L urine d<sup>-1</sup> is treated for 30 days with six fans having a power rating of 1200 W operating continuously. The energy needed for treating urine at a wastewater treatment plant is 0.9 MJ kg<sup>-1</sup>, and that for producing mineral NPK fertilizer equivalent to what is excreted in urine is 0.4 MJ kg<sup>-1</sup> (See **Supplementary Information, Supplementary Table S4**). Therefore, even at its design capacity, the pilot would have consumed ten times more energy than the conventional scenario. The energy demand is, however, comparable to what incineration toilets typically consume, 6–15 MJ kg<sup>-1</sup> urine, as they require 0.4–2 kWh flush<sup>-1</sup>.

Optimizing the energy consumption was not an objective of this study, especially since this was the first field test. The

minimum primary energy demand for evaporating water under the pilot conditions is approximately 5.3 MJ kg<sup>-1</sup> (170 kJ kg<sup>-1</sup> to increase temperature from 20 to 60°C and 2360 kJ kg<sup>-1</sup> for phase transition, and primary energy factor of 2.1). If 85% of this energy can be recovered, e.g., with a heat pump, the energy demand for drying urine can potentially be reduced to 0.8 MJ kg<sup>-1</sup>.

There are many ways to significantly reduce the energy demand of the system. The heater fans used in the pilot setup had a power rating of 700/1200 W, but could be used intermittently or substituted with ones having lower rating. The exhaust air from the system had a temperature of 35 ± 14°C but was simply vented out. Heat from the exhaust air could be recovered (around 60–95%) using air-to-air heat exchangers or heat pumps.

## Substances of Concern

One concern against the use of urine as fertilizer is its high salt concentration, since its repeated application may cause soil salinization (Mnkeni et al., 2008). The most abundant salt excreted in urine is NaCl (Putnam, 1971). This study estimated that <5% Na<sup>+</sup> and around 7% Cl<sup>-</sup> is present in the end-product obtained from alkaline urine dehydration (**Figure 2**). The Cl<sup>-</sup> concentration is similar to that in blended mineral fertilisers used in Finland and Sweden [e.g., Yara Mila 21–4–7; Yara (2020)]. The Na<sup>+</sup> concentration is high but may not be an issue as salts can normally be flushed out of soil during irrigation, rainfall events or leached by drainage (Guizani et al., 2016). However, if dried urine fertilizer is to be applied in arid regions, the Na<sup>+</sup> concentration needs to be considered.

The concentration of heavy metals excreted in urine is low and compares favorably with mineral fertilizers (Jönsson et al., 1997).



Yet, the contribution of non-essential heavy metals by the drying media can be large (Simha et al., 2020a). Of particular concern is Cd present in wood ash, as it can accumulate in food chains and risk public health (Roberts, 2014). The end-products obtained in this pilot study had an average Cd concentration of  $529 \pm 76$  mg Cd kg<sub>p</sub><sup>-1</sup>, which can be reduced to  $<370$  mg Cd kg<sub>p</sub><sup>-1</sup> by operating the system at design capacity. Yet, the Cd content is much higher than what is normally found in farmyard manure on organic cattle farms in Sweden (Jönsson et al., 2004). It is, however, less than the upcoming European Cd limits (EU, 2019) for solid organic fertilizers ( $1.5$  mg kg<sub>TS</sub><sup>-1</sup>) and organo-mineral fertilizers with  $<5\%$  P<sub>2</sub>O<sub>5</sub> ( $3$  mg kg<sub>TS</sub><sup>-1</sup>). Given these considerations, we recommend applying urine dried in wood ash only in forestry, where Cd is not a major concern (Pitman, 2006) and where the limit values are not exceeded. For agricultural applications, wood-based ash should be avoided as drying media, as several other equally effective substrates that are essentially Cd-free can be used instead (Simha et al., 2020a).

Residues of pharmaceuticals, their metabolites and transformation products are excreted via urine (Lienert et al., 2007), and there is concern against applying urine or urine-based fertilizers in agriculture (Winker, 2009). Yet, wastewater treatment plants only partially eliminate these compounds, as evidenced by their detection in urban drinking water reservoirs (Levén et al., 2016). The current best practice is to apply pharmaceutical residues in soil where there is higher likelihood for their degradation (Viskari et al., 2018; Simha et al., 2020b). In the current study, we detected very few (8 out of 126) pharmaceutical compounds in the urine and their concentrations after urine dehydration were low. This is to be expected since the users of our pilot-scale system were mainly healthy young adults undergoing military training. However, the urine was sampled only once during the study which is why some compounds (e.g., bisoprolol) were measured in the end-product but were not detected in the urine. Also, the concentration of some compounds (ibuprofen and xylometazoline) was higher than their simulated concentration, which was calculated by assuming that no degradation occurred during dehydration. The rest of the pharmaceuticals were either below the LoD or below their simulated concentration. This indicated that they were either difficult to extract from the alkaline substrate or that they were degraded during treatment. The fate of pharmaceuticals following alkaline urine dehydration is not well known and is based on previous studies on biodegradability of pharmaceuticals in water and soil, e.g., Barra Caracciolo et al. (2015). Further research is needed to understand the fate and degradation of pharmaceuticals in different matrixes in the environment.

## CONCLUSION AND IMPLEMENTATION OUTLOOK

There were several positives from this pilot-scale study. To begin with, we managed to design and engineer an on-site urine treatment system with capacity to dry 30–40 L urine d<sup>-1</sup> m<sup>-2</sup>, replicating results from earlier laboratory-scale experiments (Simha et al., 2020a). While the recovery of N was less than

what we expected, our modeling showed that the technology has potential to recover the majority of urine's nutrients and yield end-products with high fertilizer value (13.2–2.3–6.0, N-P-K as% of TS), if it is operated at full capacity. The pilot's energy demand was clearly higher than the conventional scenario, but by recovering energy (e.g., with a heat pump) it may be possible for the treatment to be energetically more favorable. This field study demonstrated that alkaline dehydration of urine is scalable and that with further research and improvements, it can be implemented in practice. An advantage of alkaline dehydration is that it is modular. Our study built three drying modules, each with dimensions of 580 × 450 × 300 mm and capacity to treat 10 L urine d<sup>-1</sup>. Simply by interlinking the modules and duplicating them, larger volumes of urine can be treated and the technology can be scaled up.

Results of our case study in Finland have a much broader context and may find use in other settings. This could be other military training sites; for instance, at the thirteen UN peacekeeping camps across the globe, where more than 100,000 UN peacekeepers (troops, police, and civilian personnel) have been deployed (United Nations Peacekeeping, 2020). Here, there is an annual potential of collecting 400 t N and 18 t P as dry concentrated urine-based fertilizer. Shelters that are managed by the International Federation of Red Cross and Red Crescent Societies following natural disasters, or shelter clusters coordinated by the UN Refugee Agency may also be appropriate settings for implementing on-site urine dehydration treatment. In 2018 alone, more than 70 million people (of which there were nearly 26 million refugees) were forcibly displaced (UNHCR, 2019) and needed shelter or various forms of shelter assistance. The need for developing better approaches to manage and recycle toilet wastes in such settings and circumstances is therefore large and urgent.

## DATA AVAILABILITY STATEMENT

The raw data supporting the conclusions of this article will be made available by the authors, without undue reservation.

## AUTHOR CONTRIBUTIONS

All authors contributed to the study conceptualization and design, manuscript revision, read and approved the submitted version. PS and CK constructed and installed the pilot-scale system. PS, CK, and E-LV performed monitoring, data collection, and formal analysis with inputs from BV and RM. BV and E-LV supervised the study, provided resources and acquired funding. PS wrote the first draft of the manuscript.

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## SUPPLEMENTARY MATERIAL

The Supplementary Material for this article can be found online at: <https://www.frontiersin.org/articles/10.3389/fenvs.2020.570637/full#supplementary-material>

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**Conflict of Interest:** BV and PS are owners of Sanitation360 AB but receive no salary.

The remaining authors declare that the research was conducted in the absence of any commercial or financial relationships that could be construed as a potential conflict of interest.

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